

Ions That Form <i>Soluble</i> Compounds	Exceptions
Li^+ Na^+ K^+	
NH_4^+	
NO_3^- ClO_4^-	

Table H
Vapor Pressure of Four Liquids

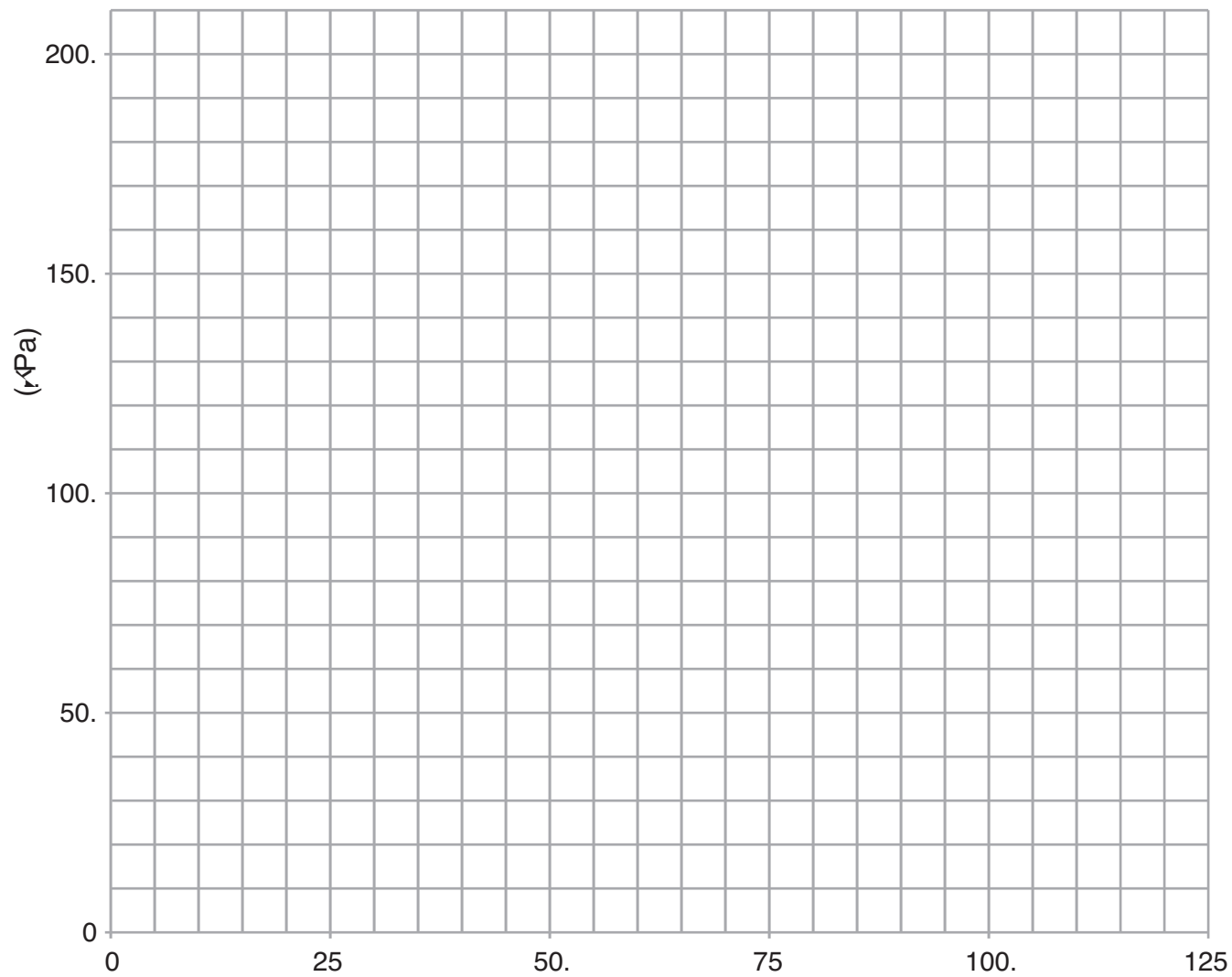


Table I
Heats of Reaction at 101.3 kPa and 298 K

Reaction	ΔH (kJ)*
<div style="text-align: center;">  </div>	

Table J
Activity Series**




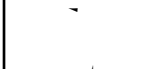
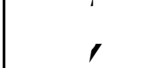

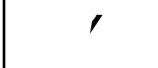













<div style="text-align: center;">  </div>
<div style="text-align: center;">  </div>
<div style="text-align: center;">  </div>
<div style="text-align: center;">  </div>
<div style="text-align: center;">  </div>
<div style="text-align: center;">  </div>
<div style="text-align: center;">  </div>
<div style="text-align: center;">  </div>
<div style="text-align: center;">  </div>
<div style="text-align: center;">  </div>
<div style="text-align: center;">  </div>
<div style="text-align: center;">  </div>
<div style="text-align: center;">  </div>
<div style="text-align: center;">  </div>
<div style="text-align: center;">  </div>
<div style="text-align: center;">  </div>
<div style="text-align: center;">  </div>
<div style="text-align: center;">  </div>
<div style="text-align: center;">  </div>
<div style="text-align: center;">  </div>

Table K
Common Acids

Table N

Table L
Common Bases

Table O

Name	General Formula	Examples	
		Name	Structural Formula
/ / / /		/ / / /	
/ / / /		/ / / /	
/ / / /		/ / / /	

Table R
Organic Functional Groups

Class of Compound	Functional Group	General Formula	Example
		R	
		R	
		R	
		R R'	
		R R'	

ft

4.00260 0
2 2

A I Ma → 12.011
S →

A I N → 6
E C I a i → 2-4

← S O i a i S a
R a w a i a a a
12C=12(a)

t : N i a
a a i i

	1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18	
1	1.00794 +1 -1	9.01218 +2	44.9559 +2	47.867 +3	50.9415 +2 +3 +4	51.966 +2 +3 +4 +5	54.9380 +2 +3 +4 +6	55.845 +2 +3 +4 +7	58.932 +2 +3	58.683 +2 +3	63.546 +2 +3	65.409 +2	69.723 +3	72.64 +2 +4	74.9216 +2 +3 +5	78.96 -2 +4 +6	79.904 -1 +1 +5	83.788 -1 +1 +5	20.180 -1
2	3 2-1	4 2-2	21 2-8-9-2	22 2-8-10-2	23 2-8-11-2	24 2-8-13-1	25 2-8-13-2	26 2-8-14-2	27 2-8-15-2	28 2-8-16-2	29 2-8-18-1	30 2-8-18-2	31 2-8-18-3	32 2-8-18-4	33 2-8-18-5	34 2-8-18-6	35 2-8-18-7	36 2-8-18-8	10 2-8
3	11 2-8-1	12 2-8-2	88.9059 +2	91.224 +3	92.9064 +2 +4	95.94 +3 +5	98 +6	101.07 +4 +6	102.906 +3	106.42 +3	107.868 +2 +4	112.41 +2	114.818 +3	118.71 +2 +4	121.760 +2 +3 +5	127.60 -3 +3 +5	128.904 -2 +4 +6	131.29 -1 +1 +5	39.948 -1 +1 +5
4	19 2-8-8-1	20 2-8-8-2	39.0983 +1	40.08 +2	41 2-8-18-12-1	42 2-8-18-13-1	43 2-8-18-13-2	44 2-8-18-15-1	45 2-8-18-16-1	46 2-8-18-18	47 2-8-18-18-1	48 2-8-18-18-2	49 2-8-18-18-3	50 2-8-18-18-4	51 2-8-18-18-5	52 2-8-18-18-6	53 2-8-18-18-7	54 2-8-18-18-8	36 2-8-18-8
5	37 2-8-18-8-1	38 2-8-18-8-2	132.905 +1	137.33 +2	180.948 +4 +5	183.84 +5	186.207 +4 +6 +7	190.23 +3 +4	192.217 +3 +4	195.08 +2 +4	196.867 +1 +3	200.59 +1 +2	204.383 +1 +2	207.2 +2 +4	208.980 +2 +4	209 +3 +5	216.904 -2 +4 +6	217.9 -1 +1 +5	218.8 -1 +1 +5
6	55 2-8-18-18-8-1	56 2-8-18-18-8-2	85.4678 +1	87.62 +2	73 -18-32-11-2	74 -18-32-12-2	75 -18-32-13-2	76 -18-32-14-2	77 -18-32-15-2	78 -18-32-17-1	79 -18-32-18-1	80 -18-32-18-2	81 -18-32-18-3	82 -18-32-18-4	83 -18-32-18-5	84 -18-32-18-6	85 -18-32-18-7	86 -18-32-18-8	54 2-8-18-18-8
7	7 -18-32-18-8-1	8 -18-32-18-8-2	223 +1	(225) +2	(262) +4	(266) +3	(272) +3	(277) +3	(276) +3	(281) +3	(280) +3	(285) +3	(284) +3	(289) +3	(288) +3	(292) +3	(?) +3	(294) +3	(222) 0
8	104 -18-32-18-9-2	105 -18-32-18-9-2	106	107	108	109	110	111	112	113**	114	115	116	117	118	119	120	121	122
9	5	6	7	8	9	10	11	12	13	14	15	16	17	18	19	20	21	22	23
10	140.116 +3 +4	140.908 +3 +4	144.24 +3	150.36 +3	151.964 +2 +3	157.25 +2 +3	158.925 +3	162.500 +3	164.830 +3	167.259 +3	168.934 +3	173.04 +2 +3	173.04 +2 +3	177.04 +2 +3	177.04 +2 +3	177.04 +2 +3	177.04 +2 +3	177.04 +2 +3	177.04 +2 +3
11	5	6	7	8	9	10	11	12	13	14	15	16	17	18	19	20	21	22	23
12	232.038 +4	231.036 +4	238.029 +4 +5	244 +4 +5 +6	244 +3 +4 +5 +6	247 +3 +4 +5 +6	247 +3 +4	251 +3 +4	252 +3	252 +3	252 +3	257 +3	258 +3	259 +2 +3	259 +2 +3	259 +2 +3	259 +2 +3	259 +2 +3	259 +2 +3
13	90	91	92	93	94	95	96	97	98	99	100	101	102	103	104	105	106	107	108
14	140.116 +3 +4	140.908 +3 +4	144.24 +3	150.36 +3	151.964 +2 +3	157.25 +2 +3	158.925 +3	162.500 +3	164.830 +3	167.259 +3	168.934 +3	173.04 +2 +3	173.04 +2 +3	177.04 +2 +3	177.04 +2 +3	177.04 +2 +3	177.04 +2 +3	177.04 +2 +3	177.04 +2 +3
15	5	6	7	8	9	10	11	12	13	14	15	16	17	18	19	20	21	22	23
16	232.038 +4	231.036 +4	238.029 +4 +5	244 +4 +5 +6	244 +3 +4 +5 +6	247 +3 +4 +5 +6	247 +3 +4	251 +3 +4	252 +3	252 +3	252 +3	257 +3	258 +3	259 +2 +3	259 +2 +3	259 +2 +3	259 +2 +3	259 +2 +3	259 +2 +3
17	90	91	92	93	94	95	96	97	98	99	100	101	102	103	104	105	106	107	108
18	140.116 +3 +4	140.908 +3 +4	144.24 +3	150.36 +3	151.964 +2 +3	157.25 +2 +3	158.925 +3	162.500 +3	164.830 +3	167.259 +3	168.934 +3	173.04 +2 +3	173.04 +2 +3	177.04 +2 +3	177.04 +2 +3	177.04 +2 +3	177.04 +2 +3	177.04 +2 +3	177.04 +2 +3

* ...

** ...

S : C H N C m p p , 91 , 2010 2011, CRC P

Table S
Properties of Selected Elements

Atomic Number	Symbol	Name	First Ionization Energy	Electro-negativity	Melting Point	Boiling Point	Density**	Atomic Radius
1								
2								
3								
4								
5								
6								
7								
8								
9								
10								
11								
12								
13								
14								
15								
16								
17								

Atomic Number	Symbol	Name	First Ionization Energy	Electro-negativity	Melting Point	Boiling* Point	Density**	Atomic Radius
41								
42								
43								
44								
45								
46								
47								
48								
49								
50								
51								
52								
53								
54								
55								
56								
57								
Elements 58–71 have been omitted.								
72								
73								
74								
75								
76								
77								
78								
79								
80								
81								
82								
83								
84								
85								
86								
87								
88								
89								
Elements 90 and above have been omitted.								

* Values are in kJ/mol
 ** Values are in g/cm³
 Source: CRC Handbook of Chemistry and Physics, 91st Edition, 2010-2011, CRC Press

Table T
Important Formulas and Equations

Density	$\rho = \frac{m}{V}$
Mole Calculations	$n = \frac{m}{M}$
Percent Error	$\% \text{ Error} = \frac{ \text{Experimental} - \text{Theoretical} }{\text{Theoretical}} \times 100$
Percent Composition	$\% \text{ Composition} = \frac{\text{mass of element}}{\text{total mass}} \times 100$
Concentration	$M = \frac{n}{V}$
	$m\% = \frac{m_{\text{solute}}}{m_{\text{solution}}} \times 100$
Combined Gas Law	$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$
Titration	$M_A V_A = M_B V_B$
Heat	$Q = C \Delta T$ $H = C m \Delta T$ $H = H_{\text{fusion}} + H_{\text{vaporization}}$
Temperature	$T_{\text{C}} = \frac{5}{9}(T_{\text{F}} - 32)$